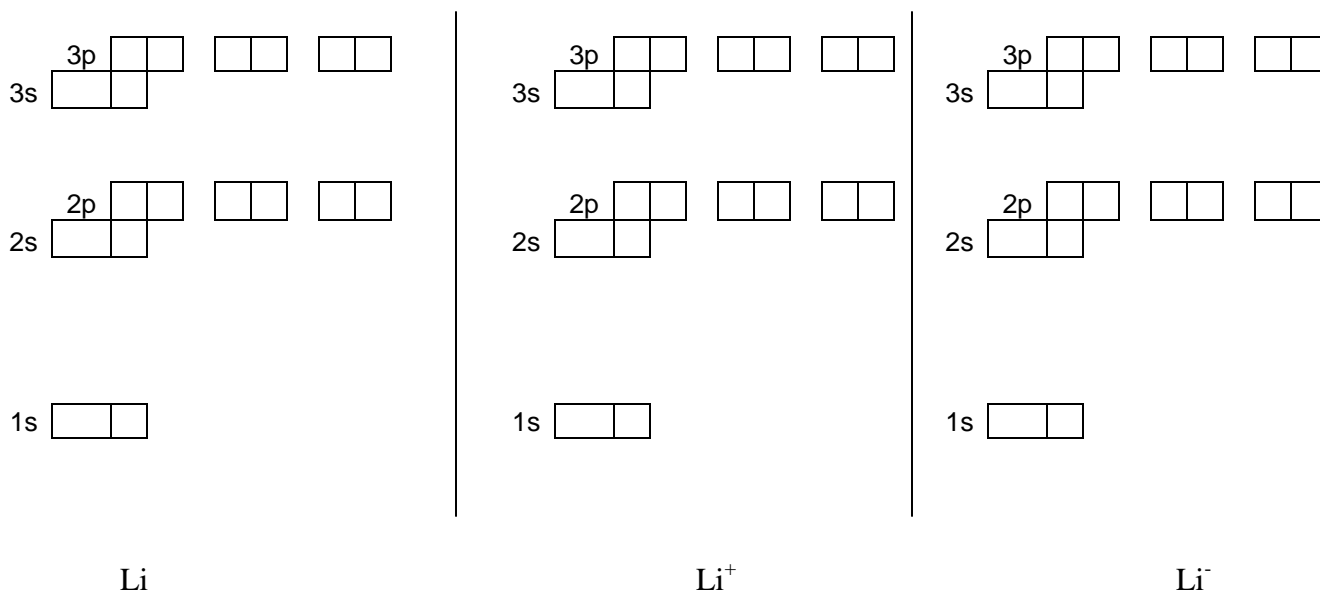


Video Notes:

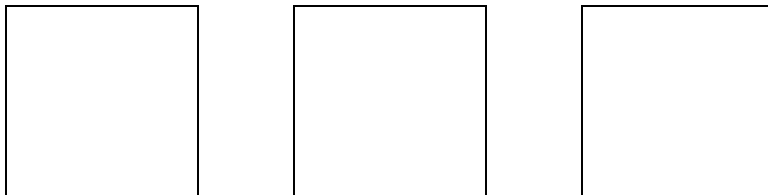
1. What are ions?
2. What are isotopes?
3. Forces that hold nucleus together are called?
4. Hydrogen is believed to be the

Section 3.2 (p81-84)

5. Define ionization:
6. Define ion:
7. Define cation:
8. Define anion:
9. Using the electron configuration “building model” demonstrate Li, Li⁺, and Li⁻.



10. Define Mass number:
11. Define atomic #:
12. Define isotope:
13. The boxes below represent the nucleus of a hydrogen atom. Add the correct number of protons and neutrons to demonstrate the different isotopes of hydrogen. Disregard electrons. Use protons = + and neutrons = n. See top of page 83.



14. Of the above isotopes, which one is the least stable? What is the ratio of protium to deuterium in the Sun (pg83)?
15. Cl has an atomic # of 17. What does this mean? _____ Cl-35 has a mass # of 35, how many neutrons does it have? _____. Cl-37 has a mass number of 37, how many neutrons does it have? _____