

Name _____

<http://www.ptable.com/> for help

Orbital and Electron Capacity for the sublevels

Sublevel	# of Orbitals	Max e-
s	1	2
p	3	6
d	5	10
f	7	14

Orbital and electron capacity

Principal energy level (n)	Type of sublevel	# of orbitals per level (n ²)	Max # of e ⁻ (2n ²)
1	s		
2	s		
	p		
3	s		
	p		
	d		
4	s		
	p		
	d		
	f		

Periodic Table of Elements

* Lanthanide Series

58	59	60	61	62	63	64	65	66	67	68	69	70	71
Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu

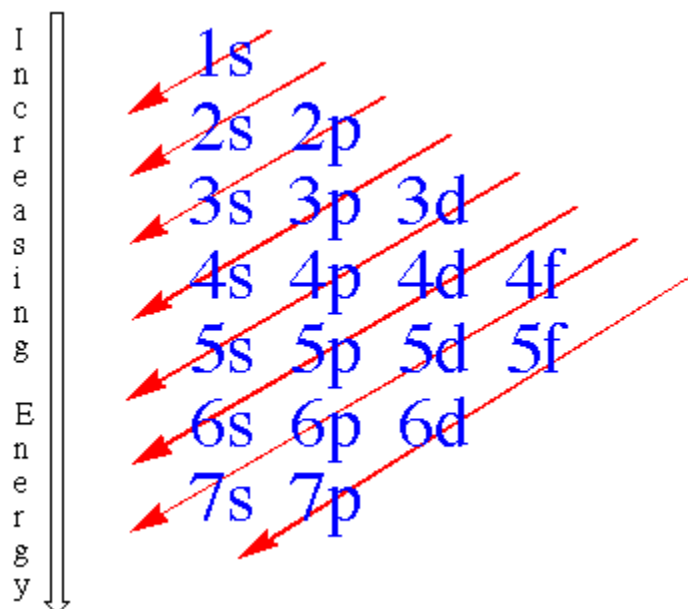
+ Actinide Series

90	91	92	93	94	95	96	97	98	99	100	101	102	103
Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

Legend - click to find out more...

- H - gas
- Li - solid
- Br - liquid
- Tc - synthetic
- Non-Metals
- Transition Metals
- Rare Earth Metals
- Halogens
- Alkali Metals
- Alkali Earth Metals
- Other Metals
- Inert Elements

Determine the electron configuration and the number of valence electrons for the following elements.



1. Lithium
2. Mg
3. Aluminum
4. Argon
5. Helium

